

START OF TERM 1 NOTES

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3.1.8.1	Born -haber cycles
3.1.8.2	Gibbs-free energy
3.1.8.3	Entropy changes
3.1.9	Rate equations
3.1.9.1	Rate equations
3.1.9.2	Determination of rate equations
3.1.10.	Equilibrium constant for homogenous systems
3.1.11	Electrode potentials and electrochemical cells
3.1.11.1	Electrode potentials and cells
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	Bronsted-Lowry acid-base equilibria in aqueous solutions
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3.1.12.2	Definition and determination of Ph
3.1.12.3	The ionic product of water
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3.1.12.5	pH curves, titrations and indicators
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3.2.5	Transition metals
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3.2.5.2	Substitution reactions
3.2.5.3	Shapes of complex ions
3.2.5.4	Formation of coloured ions
3.2.5.5	Variable oxidation states
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3.2.6	Reactions of ions in aqueous solutions
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3.3.9.1	Carboxylic acids and esters
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3.3.12.1	Condensation polymers
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3.3.13	Amino acids, proteins and DNA
3.3.13.1	Amino acids

- 3.3.13.2 Proteins
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- 3.3.13.4 DNA
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- 3.3.14** Organic synthesis
- 3.3.15** NMR
- 3.3.16** Chromatography